

Chapter 19 Study Guide - Chemical Bonding

1. What happens to the number of protons during bonding?
2. What happens to an atom for it to become an ion with a 2- charge?
3. What type of bonding occurs due to the attraction of oppositely charged ions?
4. How many valence electrons are in groups 1, 2, 13, 14, 15, 16, 17, 18?
5. Which type of bonds have properties of malleability and ductility?
6. How do atoms get a filled valence electron level?
7. What is the 2-8-8 rule?
8. Electrons in which energy level are usually involved in bonding?
9. Why don't atoms have a charge?
10. What determines how an atom will bond?
11. What are three properties of metallic bonds?
12. What type of ion is formed when atoms lose electrons?
13. What type of element tends to lose electrons when it bonds?
14. What is the correct electron dot diagram for a chlorine atom?
15. What type of bonds are represented with electron dot diagrams?
16. Why does a sodium ion have a 1+ charge?
17. What type of bond joins diatomic molecules?
18. How do atoms bond?
19. Do nonmetals require a lot or a little energy to lose electrons?
20. What is a crystal lattice?
21. What type of bonds are in a staple?
22. What type of bonds are in table sugar?